

MATH HANDBOOK TRANSPARENCY MASTER

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Scientific Notation

Use with Appendix B,
Scientific Notation

Scientists need to express small measurements, such as the mass of the proton at the center of a hydrogen atom (0.000 000 000 000 000 000 000 001 673 kg), and large measurements, such as the temperature at the center of the Sun (15 000 000 K). To do this conveniently, they express the numerical values of small and large measurements in scientific notation, which has two parts.

A number in which only one digit is placed to the left of the decimal

$$N \times 10^n$$

An exponent of 10 by which the number is multiplied

Thus, the temperature of the Sun, 15 million kelvins, is written as 1.5×10^7 K in scientific notation.

Positive Exponents Express 1234.56 in scientific notation.

	1234.56	
Each time the decimal place is moved one place to the left,	$1234.56 \times 10^0 = 123.456 \times 10^1$	the exponent is increased by one.
	$123.456 \times 10^1 = 12.3456 \times 10^2$	
	$12.3456 \times 10^2 = 1.234\ 56 \times 10^3$	
	$1.234\ 56 \times 10^3$	

Negative Exponents Express 0.006 57 in scientific notation.

	0.006 57	
Each time the decimal place is moved one place to the right,	$0.006\ 57 \times 10^0 = 0.0657 \times 10^{-1}$	the exponent is decreased by one.
	$0.0657 \times 10^{-1} = 0.657 \times 10^{-2}$	
	$0.657 \times 10^{-2} = 6.57 \times 10^{-3}$	
	6.57×10^{-3}	

MATH HANDBOOK TRANSPARENCY WORKSHEET**1****Scientific Notation****Use with Appendix B,
Scientific Notation****1.** Express each of the following numbers in scientific notation.**a.** 230
_____**b.** 5601
_____**c.** 14 100 000
_____**d.** 56 million
_____**e.** 2/10
_____**f.** 0.450 13
_____**g.** 0.089
_____**h.** 0.000 26
_____**i.** 0.000 000 698
_____**j.** 12 thousandth
_____**2.** Express each of the following measurements in scientific notation.**a.** speed of light in a vacuum, 299 792 458 m/s
_____**b.** number of seconds in a day, 86 400 s
_____**c.** mean radius of Earth, 6378 km
_____**d.** density of oxygen gas at 0°C and pressure of 101 kPa, 0.001 42 g/mL
_____**e.** radius of an argon atom, 0.000 000 000 098 m

SCIENTIFIC NOTATION

Name _____

Scientists very often deal with very small and very large numbers, which can lead to a lot of confusion when counting zeros! We have learned to express these numbers as powers of 10.

Scientific notation takes the form of $M \times 10^n$ where $1 \leq M < 10$ and "n" represents the number of decimal places to be moved. Positive n indicates the standard form is a large number. Negative n indicates a number between zero and one.

Example 1: Convert 1,500,000 to scientific notation.

We move the decimal point so that there is only one digit to its left, a total of 6 places.

$$1,500,000 = 1.5 \times 10^6$$

Example 2: Convert 0.000025 to scientific notation.

For this, we move the decimal point 5 places to the right.

$$0.000025 = 2.5 \times 10^{-5}$$

(Note that when a number starts out less than one, the exponent is always negative.)

Convert the following to scientific notation.

1. $0.005 =$ _____

6. $0.25 =$ _____

2. $5,050 =$ _____

7. $0.025 =$ _____

3. $0.0008 =$ _____

8. $0.0025 =$ _____

4. $1,000 =$ _____

9. $500 =$ _____

5. $1,000,000 =$ _____

10. $5,000 =$ _____

Convert the following to standard notation.

1. $1.5 \times 10^3 =$ _____

6. $3.35 \times 10^{-1} =$ _____

2. $1.5 \times 10^{-3} =$ _____

7. $1.2 \times 10^{-4} =$ _____

3. $3.75 \times 10^{-2} =$ _____

8. $1 \times 10^4 =$ _____

4. $3.75 \times 10^2 =$ _____

9. $1 \times 10^{-1} =$ _____

5. $2.2 \times 10^5 =$ _____

10. $4 \times 10^0 =$ _____